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gas occupies a volume of 2.45 L at a pressure of 1.03 atm and a temperature of 293 K.

## 13-3 Practice Problems

the factory is 761.3 mm Hg. The pressure inside the bag is the same. What is the pressure inside the bag of potato Chips in Pa? The same bag of potato chips from Problem 5 is shipped to a town in Colorado, where the atmospheric pressure is 99.82 kPa. What is the difference (in Pa) between the pressure in the bag and the atmospheric

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Problems 1. A gas occupies a volume of 458-mL at a pressure of 1.01- kPa and temperature of 295-K. When the pressure is changed, the volume becomes 477-mL.

If there has been no change in temperature, what is the new pressure? 6. What will be the volume of a gas sample at 309-K if its volume at 215-K is 3.42-L?

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13.3: Factors Affecting Solubility Conceptual Problems  
Use the kinetic molecular theory of gases discussed in Chapter 10 to explain why the solubility of virtually all gases in liquids decreases with increasing temperature.

13.E: Properties of Solutions (Exercises) - Chemistry

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Practice Chemistry with Worked Chemistry Problems  
13—4 Practice Problems 1. What volume would be

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occupied by 100. g of oxygen gas at a pressure of 1.50 atm and a temperature of 250°C?  $P_1 V_1 = n R T_1$  • (IF  $T_1 = 273\text{K}$ ) 2. An air-filled balloon has a volume of 5 L at 0.94 atm and 250°C. Soon after, the pressure changes to 0.99 atm and the temperature changes to 0°C. What is the new volume of the balloon?  $V_2 = 1.26\text{ L}$  ...

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Name:\_\_\_\_\_ Date:\_\_\_\_\_ Chemistry 10-3 Practice  
Problems (Continued) 13. Find the empirical formula of  
a compound, given that the compound is found to be  
47.9 percent zinc and 52.1 percent chlorine by mass.  
20.



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