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1. In a bottle of unopened cola, the CO_2 gas dissolved in the liquid is in equilibrium with the CO_2 gas above the liquid.

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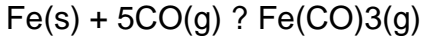
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Chemical Equilibrium

Magnesium hydroxide, Mg(OH) ...



$[\text{Fe(CO)}_3] / [\text{CO}]^5$. If the equation



$2\text{H}_2\text{(g)}$ is for a system at equilibrium, increasing the temperature will cause.

$[\text{CH}_3\text{OH}]$ to decrease and $[\text{CO}]$ and $[\text{H}_2]$ to increase.

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STUDY. PLAY. Reversible Reaction. A chemical reaction in which the products can react to reform the reactants. Chemical Equilibrium. When the rate of its forward reaction equals the rate of its reverse reaction and the concentrations of its products and reactants remain unchanged.

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1. $A + B \rightleftharpoons C + D$ (forward reaction) $C + D \rightleftharpoons A + B$ (reverse reaction)

Equilibrium (forward rate = reverse rate) remain constant. The ratio of the mathematical product $[C]^x[D]^y$ to the

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mathematical product $[A]^n \times [B]^m$ for this reaction has a definite value at a given temperature.

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“Equilibrium” is the seventh chapter in

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the NCERT class 11 chemistry textbook. Several important concepts such as equilibrium constants, buffer solutions, and the common-ion effect is explained in this chapter.

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Similarly, in Chapter 13, we discussed saturated solutions, another example of a physical equilibrium, in which the rate of dissolution of a solute is the same as the rate at which it crystallizes from solution. In this chapter, we describe the methods chemists use to quantitatively describe

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the composition of chemical systems
at equilibrium ...

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Ion Equilibria; Chapter 17 -
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Scientific study of chemistry, focusing on such topics as reaction rates.

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