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Problems - Chemistry

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$$\begin{aligned}M_1 V_1 &= M_2 V_2 \quad (1.71 \text{ M}) \\(25.0 \text{ mL}) &= M_2 (65.0 \text{ mL}) \\M_2 &= 0.658 \text{ M. } M = \text{mol/L} = \\(25.0/40.0) &/ (0.325) = 1.92 \\ \text{mol/L. } g &= (M) (L) (\text{FW}) = \\(0.400) &((0.225) (119) =\end{aligned}$$

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Answers

10.7 g. (25.0g) (1 mol/101 g) (1000mL/0.650 mol) = 381 mL. Zn (NO₃)₂ AlCl₃ CuAc₂. 2 mol Ca (OH)₂ = mol HBr 2 (g/74) = (3.00) (0.0500) 5.55 g Ca (OH)₂.

~~Molarity 1 (Worksheet)~~

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Answers

CHEMISTRY: A Study of Matter
© 2004, GPB 10.18b 5. 125 cm³ of solution contains 3.5 moles of solute. What is the molarity of the solution? ?
g KNO₃ = 0.175 mol KNO₃ ×
101.1 g KNO₃ 1 mol KNO₃ =

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Answers
17.7 g KNO_3 $M = 3.5 \text{ mol}$
0.125 L = 28 M 6. Which
solution is more
concentrated? Solution "A"
contains 50.0 g of CaCO_3 in
500.0 mL of solution.

~~Molarity: Molarity = 1. 2.~~

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Answers

What is the molarity of the following solutions given that: 1) 1.0 moles of potassium fluoride is dissolved to make 0.10 L of solution. $1.0 \text{ mole KF} = 10. \text{ M}$ 0.10 L soln 2) 1.0 grams of potassium fluoride is

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Answers dissolved to make 0.10 L of solution. 1.0 g KF x 1 mole
KF = 0.0172 mol KF 58 g KF
0.0172 mol KF = 0.17 M 0.10
L soln

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The unit usually used for molarity in chemistry is mol/L and is represented by the symbol M. Molarity is calculated by determining the number of liters of a

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Answers, determining the number of moles of solute in a solution, and then dividing the number moles of solute by the liters of solution. This customizable and printable worksheet is designed to help students

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Solution: $MV = \text{grams} / \text{molar}$
 $\text{mass. (x) (1.000 L) = 245.0}$
 $\text{g} / 98.0768 \text{ g mol}^{-1} \cdot 1. \text{ x} =$
 $2.49804235 \text{ M. to four sig}$
 $\text{figs, } 2.498 \text{ M. If the volume}$
 $\text{had been specified as } 1.00 \text{ L}$

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(as it often is in problems like this), the answer would have been 2.50 M, NOT 2.5 M.

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#1 — 10~~

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4.3 x 10²² molecules H₃PO₄ x

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moles H_3PO_4 . 6.022×10^{23}
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molecules are in 48.0 grams
of NaOH? $48.0 \text{ molecules NaOH}$
 $\times 1 \text{ mole NaOH} \times 6.022 \times 10^{23}$
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mole NaOH. $= 7.23 \times 10^{23}$

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